Knowledge that will change your world

3-11-20

Following pathways with isotopes

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Synopsis

- **Extra value in the M+2 isotope peak**
- **Using isotopes in tracing metabolic and physiologic pathways**
  - Historical
  - Fluxomics
  - Physiology
- **Isotopes and enhanced chemical detection of metabolites**
- **Disturbing energy levels**

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Value of natural isotopes

- The natural abundance of isotopes enables the investigator to determine the charge state of an ion
- The principal contribution to $[M+H]^+$ or $[M-H]^-$ isotope ions comes from ^{13}C (~1.1% of all carbon atoms)
- The intensity of the $[M+H]^+$ or $[M-H]^-$ ^{13}C isotope ion increases relative to the number of carbon atoms
- There is often an observable $^{13}C_2$ isotope peak

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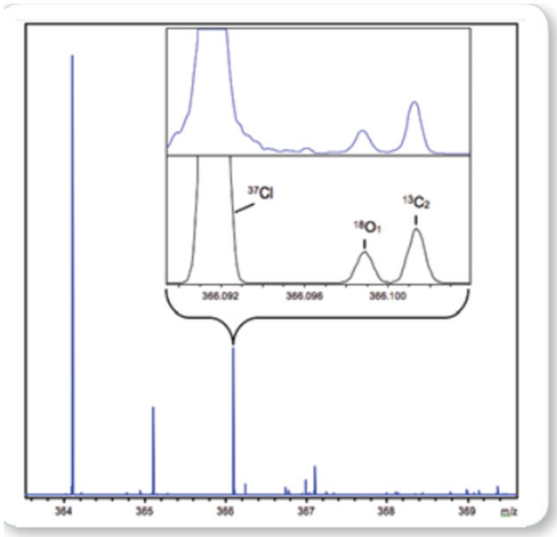
The importance of the M+2 ion



From Bruker

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Value of the [M+2] peak

- The mass difference due to a nominal increase in mass of 2 contains a lot of information
  - These are isotopic mass differences for each of the common elements
    - $^{1}H_2/^{2}H_2$ 2 x 1.006277 = 2.012554 (0.012%)
    - $^{12}C_2/^{13}C_2$ 2 x 1.003355 = 2.006710 (1.078%)
    - $^{14}N_2/^{15}N_2$ 2 x 0.997035 = 1.994079 (0.364%)
    - $^{16}O_2/^{17}O_2$ 2 x 1.004217 = 2.008434 (0.038%)
    - $^{16}O_2/^{18}O_1$ 1 x 2.004246 = 2.004246 (0.205%)
    - $^{32}S_2/^{33}S_2$ 2 x 0.999387 = 1.998774 (0.752%)
    - $^{32}S_2/^{34}S_1$ 1 x 1.995796 = 1.995796 (4.252%)
- Needs the highest possible mass resolution
  - FT-ICR

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Using isotopes to trace a pathway

- Early studies (1930s) used ^{2}H, ^{13}C and ^{15}N labeling to map pathways
  - Limited to 1-200 *m/z* mass range
- 1950s/60s ^{14}C-radiotracers
  - 2D-Paper or thin layer chromatography
  - Radio gas chromatography
    - labeling of specific carbon atoms

David Rittenberg

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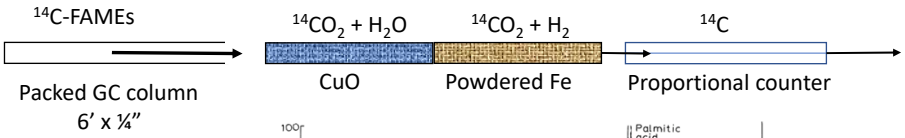
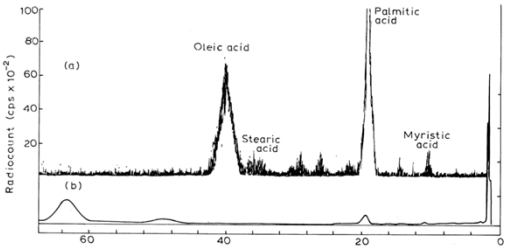
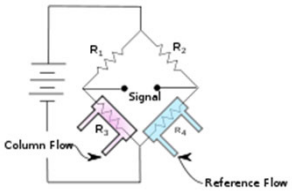
Radio-gas chromatography of FAMEs

In Tony James' lab







Katharometer to detect H_2 by thermal conductivity

We were studying an isomer of palmitoleic acid – from β-oxidation of oleic acid, or direct desaturation? Collected the 16:1 peak using a gas density balance (no degradation) in **ether-soaked cotton wool**. Subjected to oxidation with permanganate-periodate – identified C_{11} monobasic acid and C_7 dibasic acid, i.e., 16:1Δ^7

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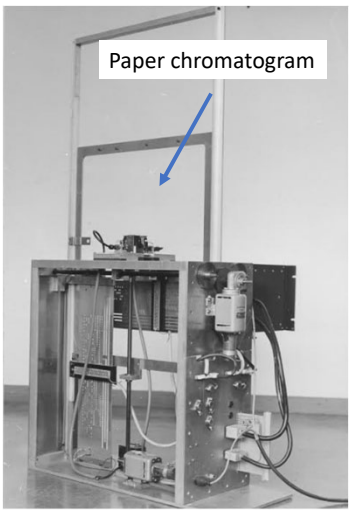
Early beginnings of metabolomics in London

- **Sir Ernst Chain (1945 Nobel Laureate – the biochemist who characterized penicillin)**
  - **Also renown for his work on microanalysis**
- **Used 2D-paper chromatography to resolve glycolytic, Krebs cycle and amino acids derived from ^{14}C-glucose**
  - **Geiger counter mounted on a typewriter frame**
  - **Digitized the collected data and prepared computer-generated figures**

METABOLOMICS

Keith Mansford



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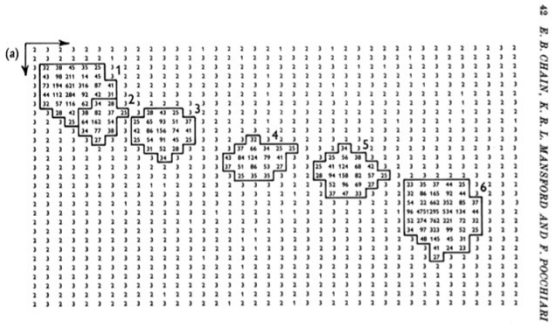
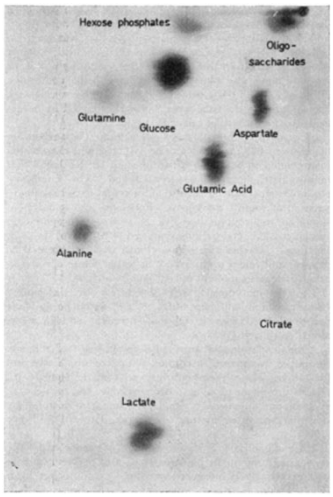
Radiochromatography examples



J Physiol (1960) 154:39
E.B. Chain, K.R.L. Mansford and F. Pocchiari



Autoradiogram of ^{14}C-glucose metabolites from an isolated perfused Langendorff rat heart preparation. The metabolites were separated by 2D-paper chromatography.

The conditions were:
1st dimension: butan-l-ol-acetic acid-water (40:11:25, by vol.) for 16hr.;
2nd dimension: (-) phenol-aq. NH_3 (sp.gr. 0.88)-water (80:1:20, by vol.) for 24hr.

Biochem. J. (1969) 115, 537
E.B. Chain, K.R.L. Mansford and L.H. Opie

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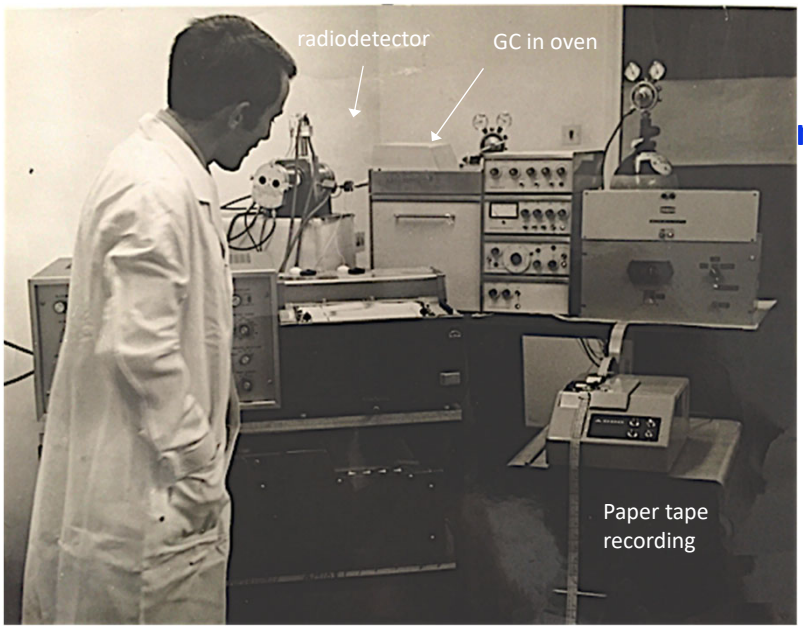

Radio-GC analysis
metabolomics in its infancy

Radio gas-liquid chromatography with digitization of collected data

Developed this for my PhD work (1967-1970) to study glucose metabolism in acellular slime mold, *Physarum polycephalum*

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Physarum polycephalum

The many headed slime mold

This is a single cell that spreads out to cover a petri dish

The cytoplasm is pushed to one end and back over a 30-45 s period

The cytoplasm has properties of a liquid and a solid

https://www.youtube.com/watch?v=Ia3kWlS_OZU

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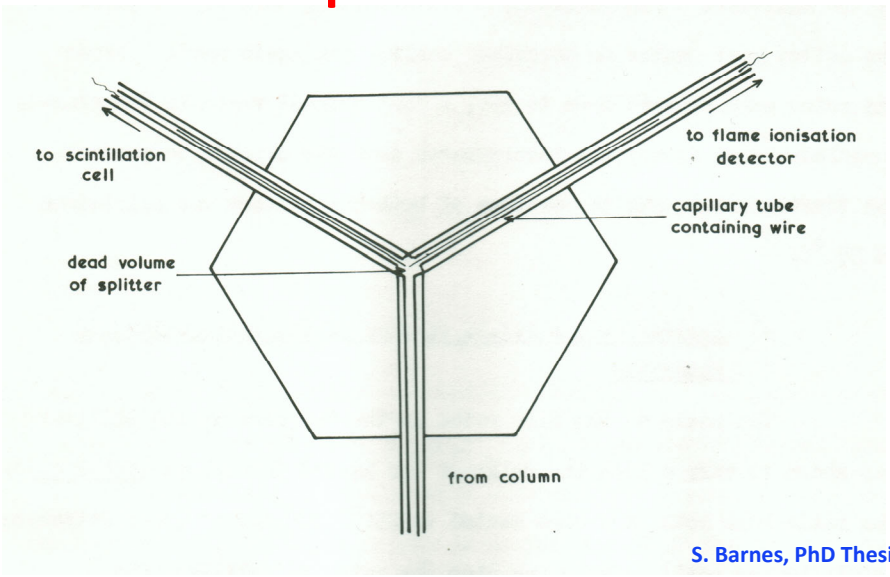
Stream splitter for radio GC



S. Barnes, PhD Thesis

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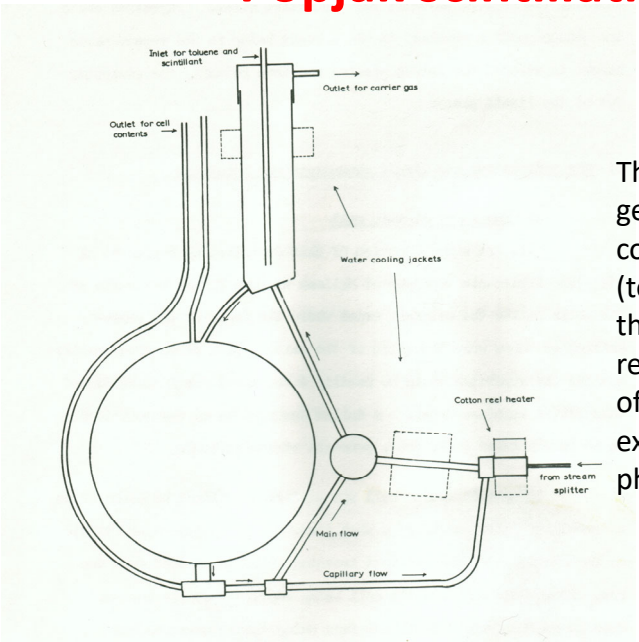
Popjak scintillation cell



The key to this device was the mixing generated by the gas from the GC column causing (scintillation) fluid (toluene) to flow out of the bottom of the scintillation chamber, both to aid recirculation and to provide a source of solvent vapors that more efficiently extracted the compounds in the gas phase.

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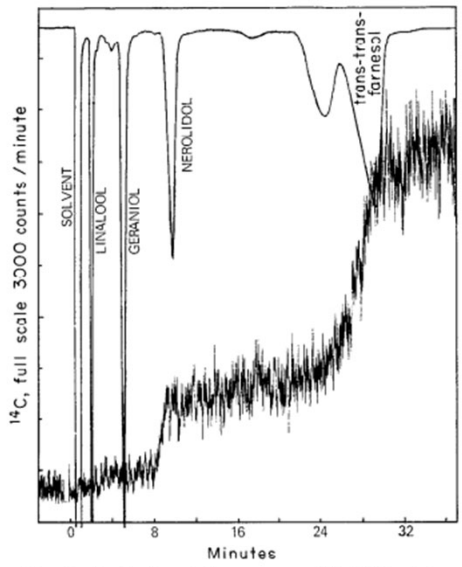

FIG. 19. Gas-liquid radiochromatogram of F_1 (*cf.* Fig. 4) from the chromatography on DEAE-cellulose of the butanol extract of an incubation of liver microsomes with farnesyl pyrophosphate. Nearly 30% of the total radioactivity in the specimen was accounted for by ^{14}C in nerolidol. Cochromatography with added markers of linolool, geraniol, nerolidol, *cis-trans*- and *trans-trans*-farnesol; simultaneous recording of mass and radioactivity detector.

Application to the discovery of a new intermediate in squalene biosynthesis

It's worth reading this 1969 article in J Biol Chem for the depth of analysis that was undertaken to prove the identity of this intermediate

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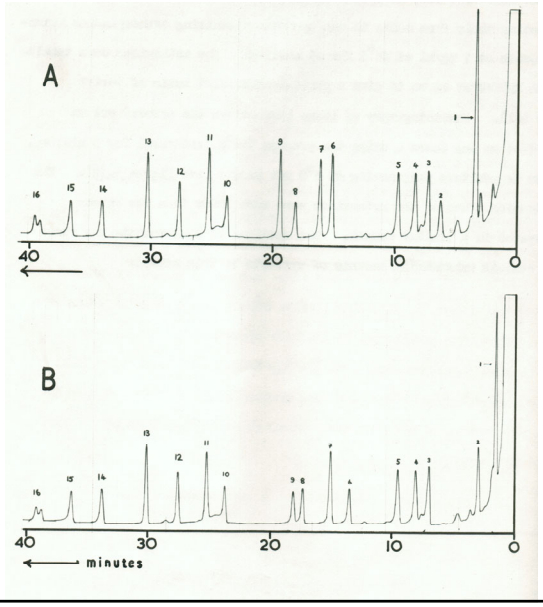
GC of glycolytic and Krebs cycle intermediates



Temperature programming of TMS ester/ethers on a 5' x ¼ inch packed column of Chromosorb W coated with OV-1 liquid phase

1=pyruvate , 2=?? , 3=phosphate , 4=succinate , 5=fumarate, 6=oxaloacetate, 7=malate, 8=αKG, 9=hexadecane, 10=αGP, 11=citrate, 12=α–D-glucose, 13=β–D-glucose, 14=docosane, 15=F6P, 16=G6P

S. Barnes, PhD Thesis

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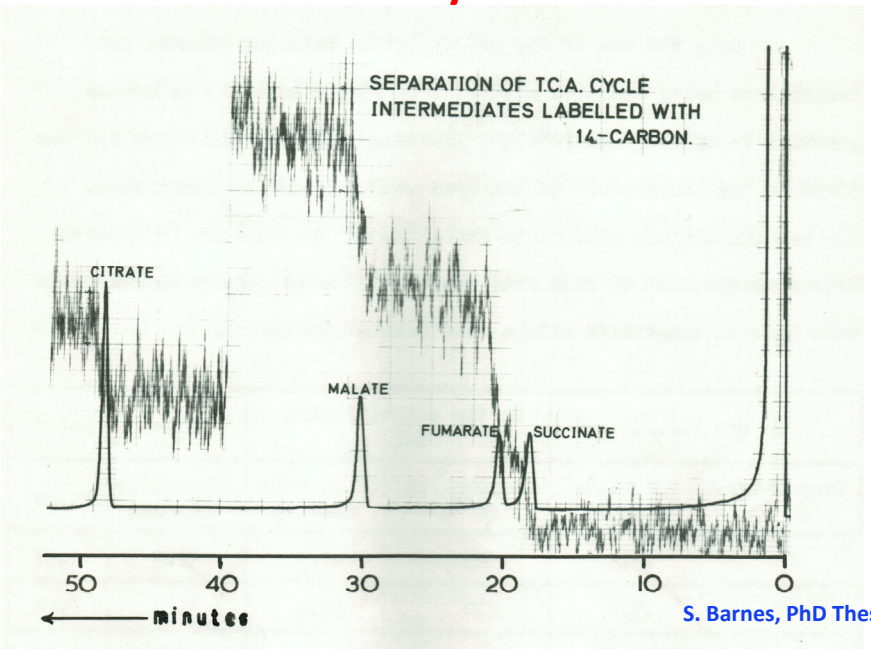
Radio-GC of Krebs Cycle intermediates



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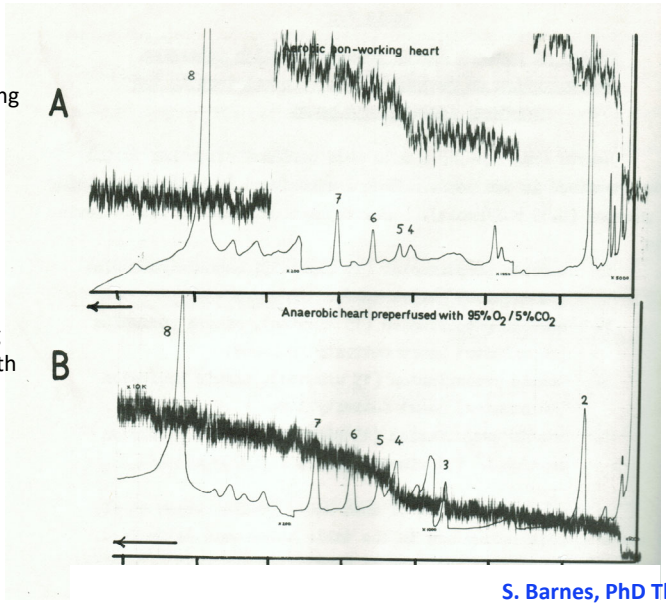
Radio GC analysis of beating heart

Aerobic non-working heart

Anaerobic beating heart perfused with 95% O_2/5% CO_2



S. Barnes, PhD Thesis

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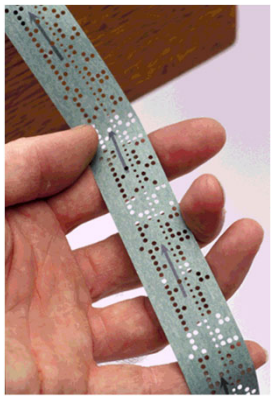
Software for data analysis on a PDP9 computer

```
DIMENSION IBUFF(1000), IDATA(725,2)
COMMON IBUFF,IDATA,ITIME,INT,ISIG
5       ITIME=0
7       CALL TAPE(ISIG,INTA)
8       IWRITE(1,1001)INTA
        INT=INTA/100
10      NPOINT=1
11      CALL TAPE(ISG,IDATA(NPOINT,1))
        IF (ISIG.EQ.1) GO TO 16
        IF (ISIG.EQ.2) GO TO 13
        WRITE(1,1001)NPOINT
GO TO 11
13      IF (IDATA(1,1).GT.940) GO TO 11
        IF (IDATA(NPOINT,1).EQ.0) GO TO 11
        IF (NPOINT.EQ.725) GO TO 16
15      NPOINT=NPOINT+1
        GO TO 11
```



Punched tape data
1 data point/sec

Digital PDP computer
Had a screen

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Fluxomics

See talk by Teresa Fan

https://www.uab.edu/proteomics/metabolomics/workshop/2018/videos/fan_day3.html

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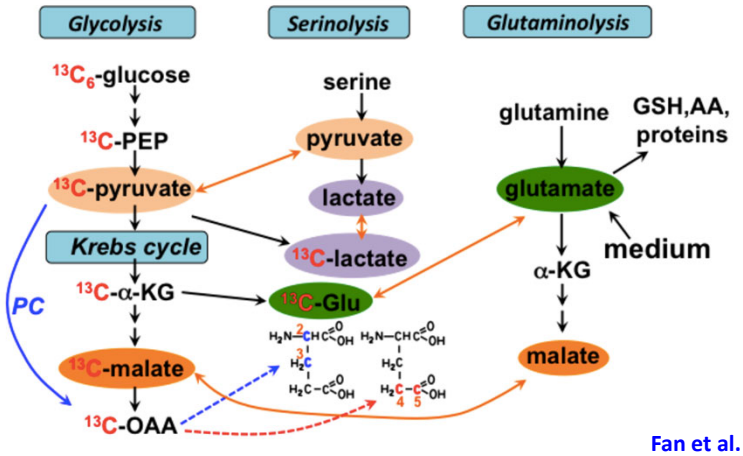
Fluxomics with stable isotopes

- **A feature of many metabolites is that they have multiple origins**



Fan et al. 2013

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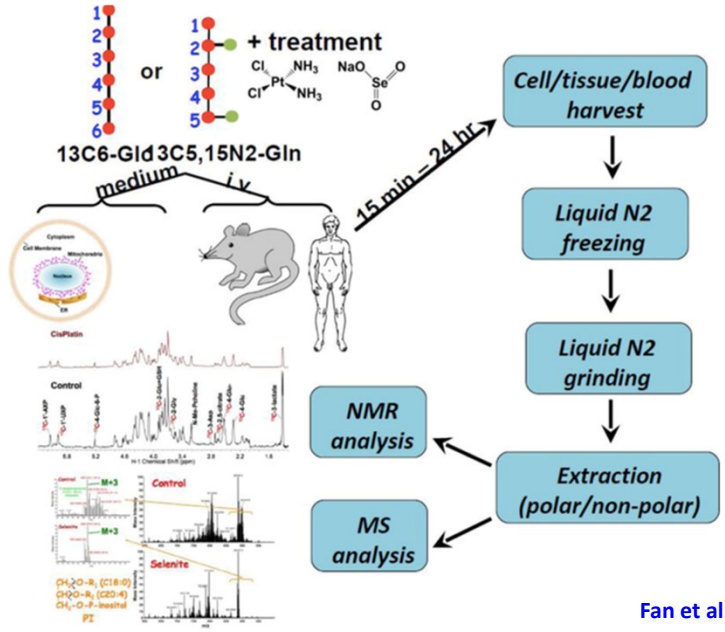
Stable isotope resolved metabolomics



Fan et al. 2013

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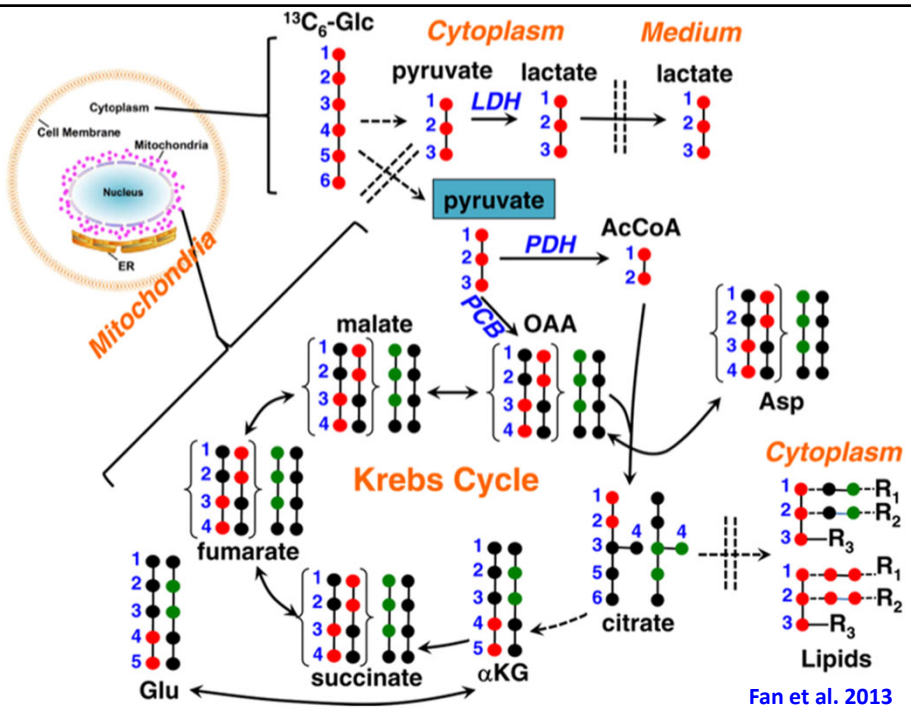

Fan et al. 2013

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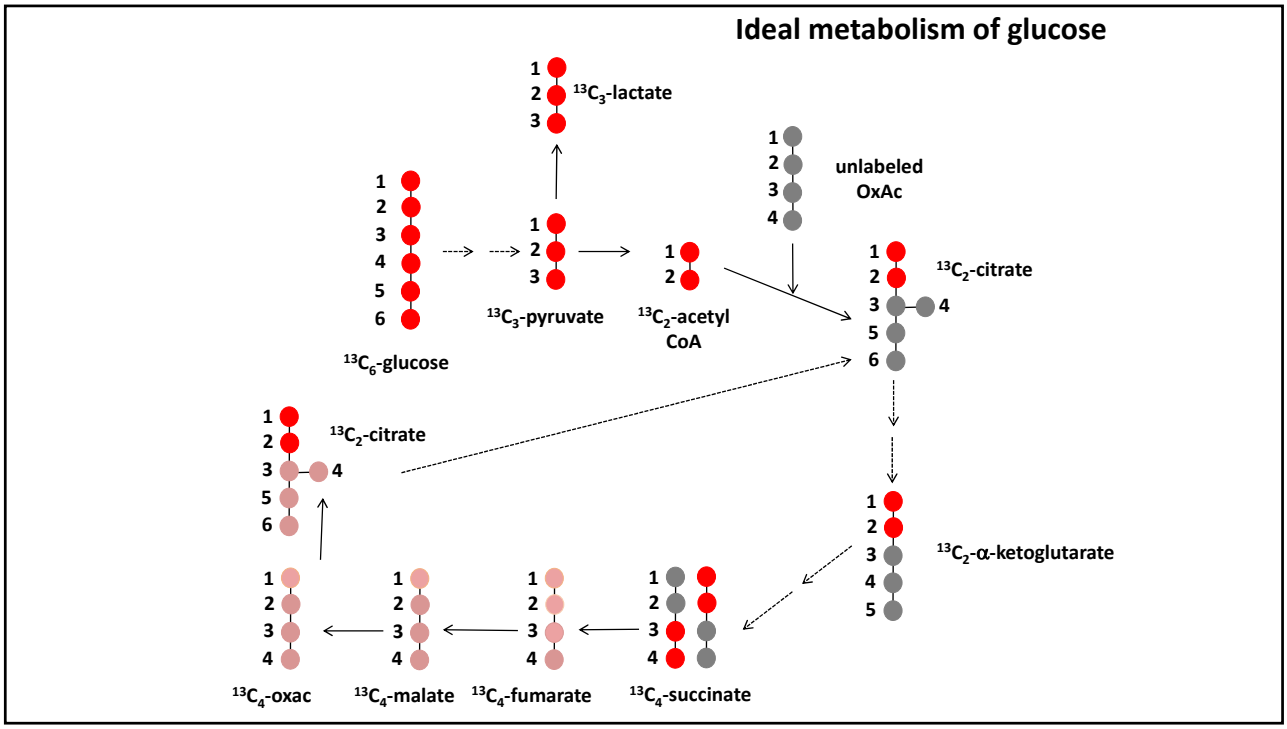
Ideal metabolism of glucose
$^{13}C_3$-lactate
unlabeled OxAc
$^{13}C_6$-glucose
$^{13}C_3$-pyruvate
$^{13}C_2$-acetyl CoA
$^{13}C_2$-citrate
$^{13}C_2$-citrate
$^{13}C_2$-α-ketoglutarate
$^{13}C_4$-oxac
$^{13}C_4$-malate
$^{13}C_4$-fumarate
$^{13}C_4$-succinate

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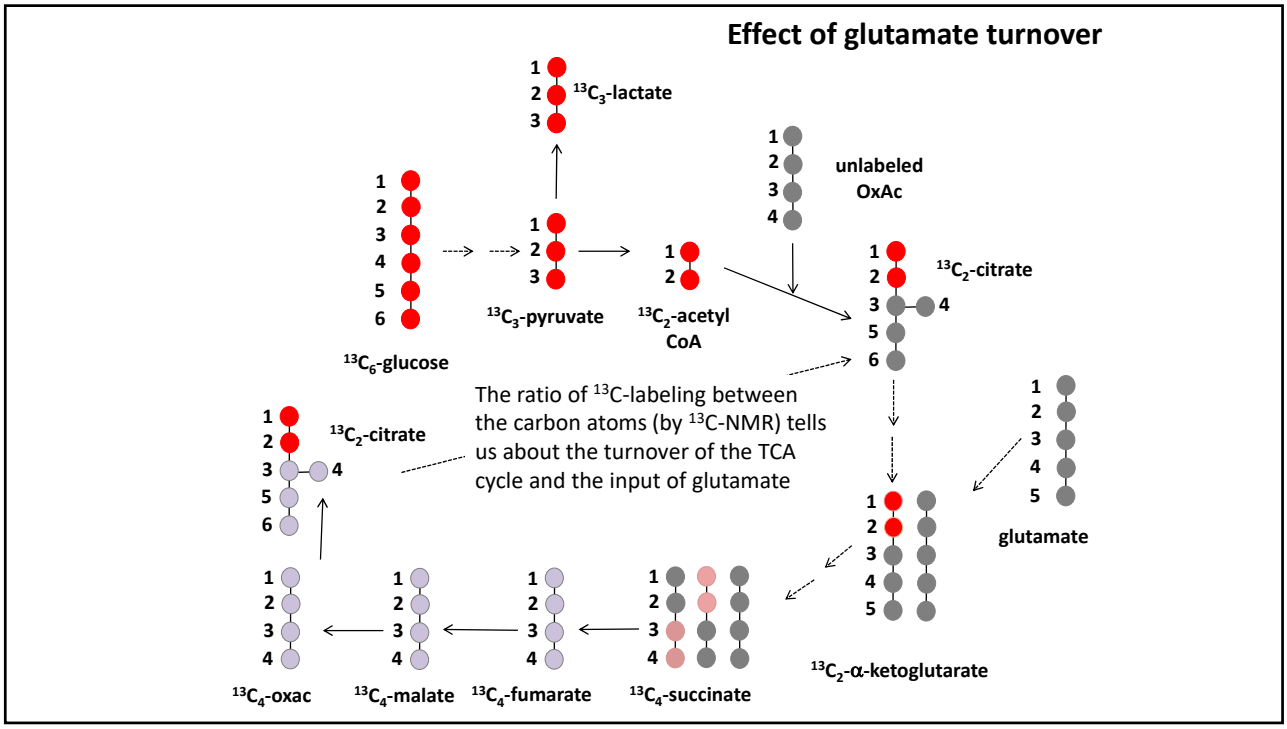
Effect of glutamate turnover
$^{13}C_3$-lactate
unlabeled OxAc
$^{13}C_6$-glucose
$^{13}C_3$-pyruvate
$^{13}C_2$-acetyl CoA
$^{13}C_2$-citrate
The ratio of ^{13}C-labeling between the carbon atoms (by ^{13}C-NMR) tells us about the turnover of the TCA cycle and the input of glutamate
$^{13}C_2$-citrate
glutamate
$^{13}C_2$-α-ketoglutarate
$^{13}C_4$-oxac
$^{13}C_4$-malate
$^{13}C_4$-fumarate
$^{13}C_4$-succinate

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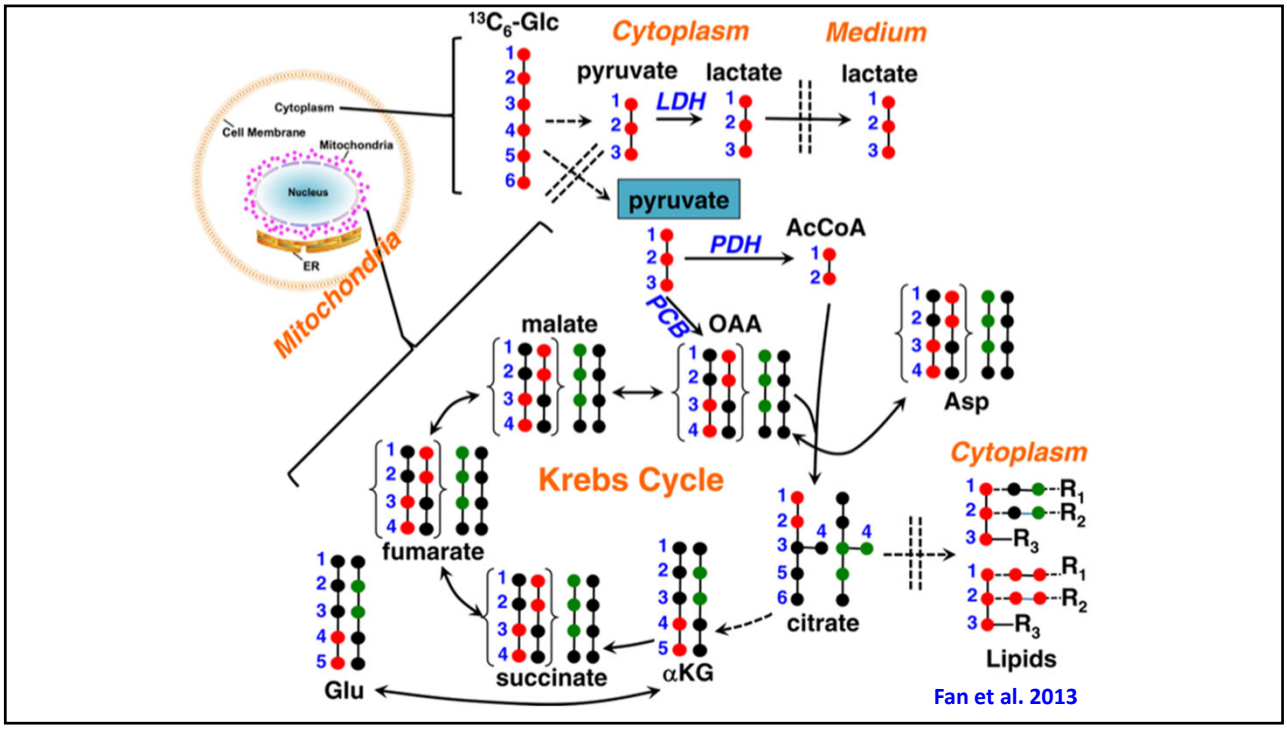

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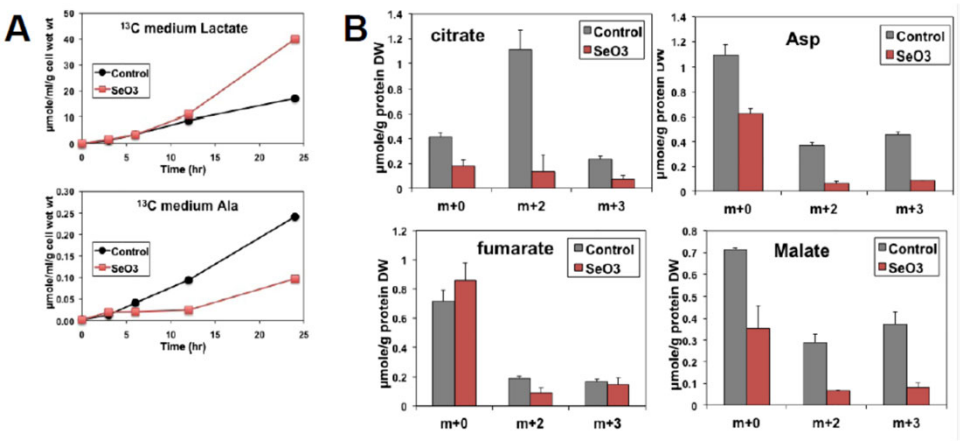
Effect of selenite on pools of intermediates



Pyruvate carboxylase converts pyruvate to oxaloacetate and by-passes the early steps in the Krebs cycle. Treatment of the cells with selenite blocks this step and the ^{13}C-content of citrate sharply decreases

Fan et al. 2013

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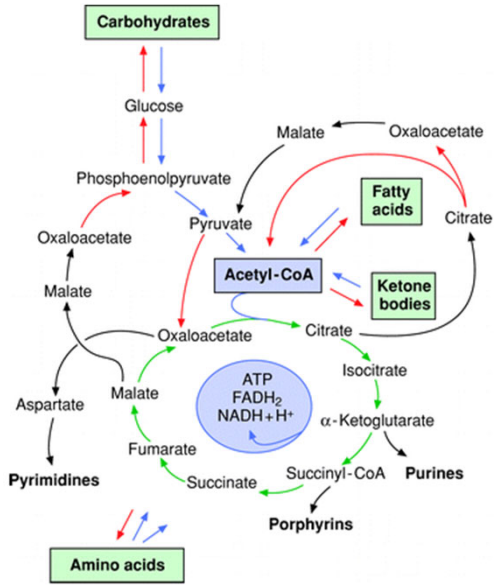
Anaplerotic reactions



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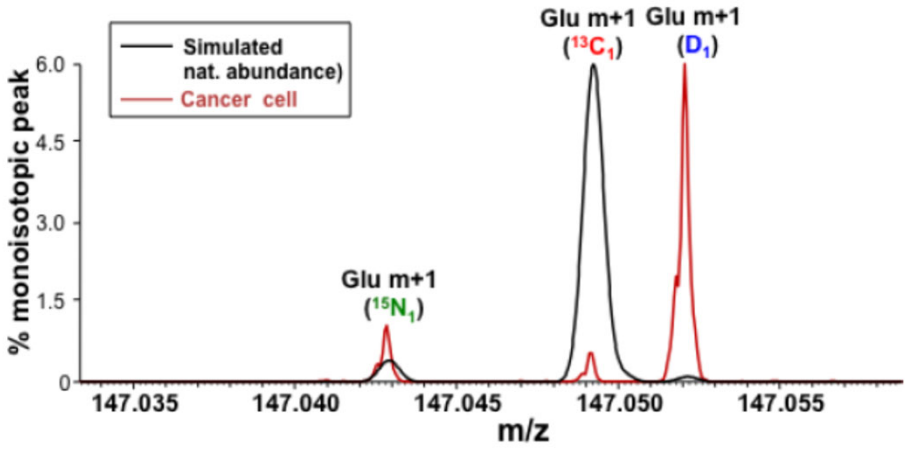
High resolution FT-ICR-MS



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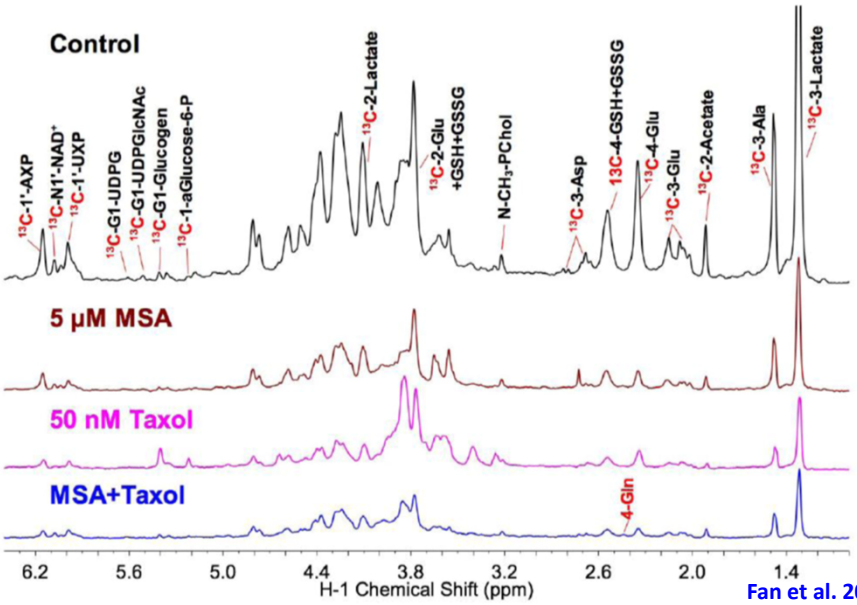
Use of ^{1}H-^{13}C-NMR



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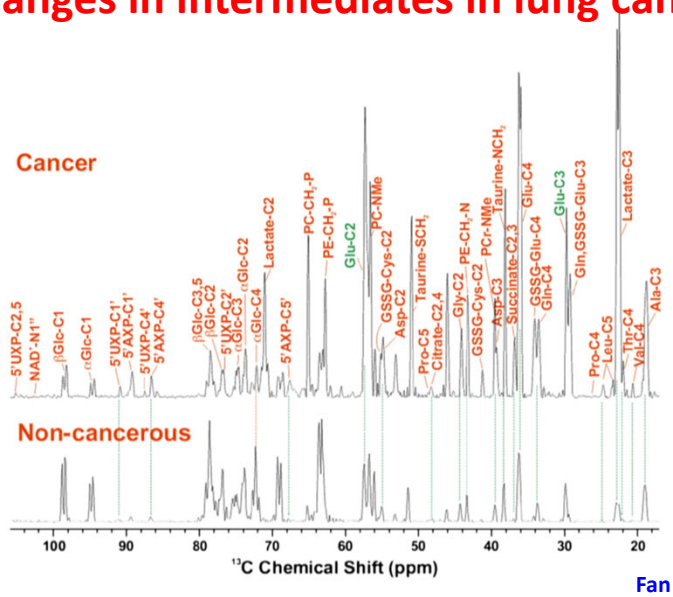
Changes in intermediates in lung cancer



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Biological NMR

- If ^{13}C-labeled precursors are used, there is a very much enhanced set of ^{13}C NMR resonances
- You have a choice between analysis of a biological extract (have all the time you need)
- And direct analysis in tissue:
  - Surface coil technology in the living animal
  - Magic Angle Spinning on a piece of tissue

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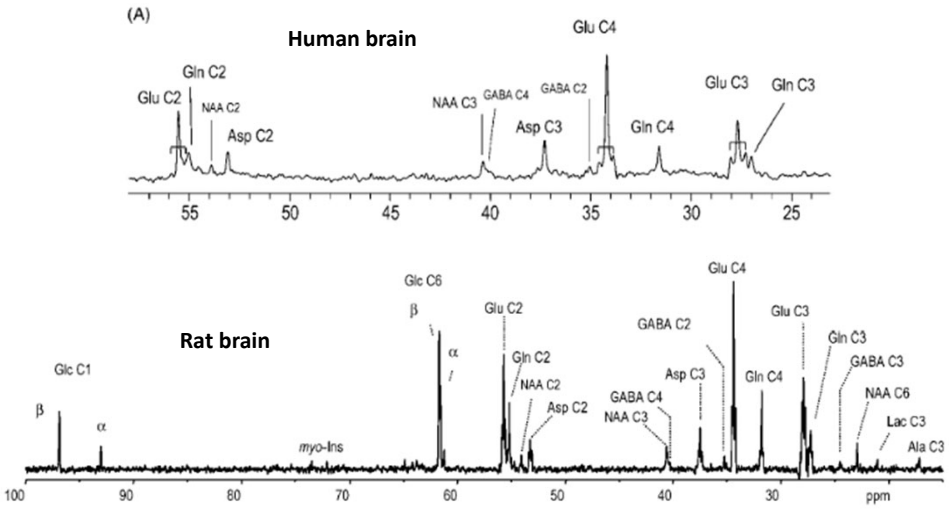
NMR analysis of metabolites from ^{13}C-labeled precursors using pulse sequences



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Probing the depths of metabolite penetration into tissues

Ultimate sensitivity by sacrificing metabolite identity at physiologic sites by ^{14}C-labeling the precursor of interest

Technology the same as the one used for radiocarbon dating

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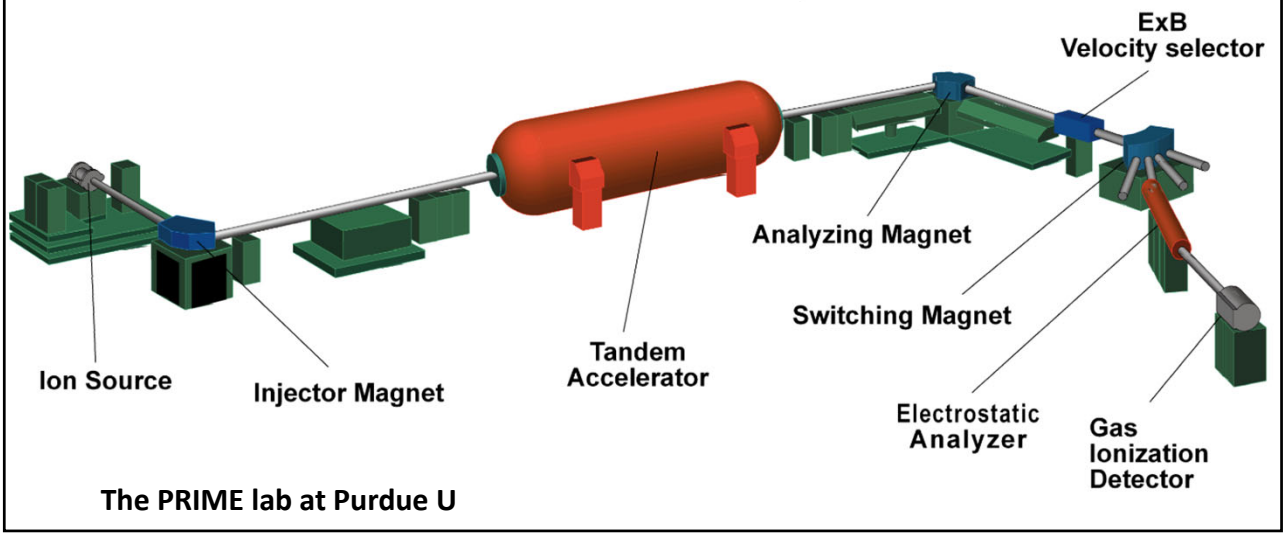
Accelerator mass spectrometry (AMS)

The ultimate mass spectrometer



The PRIME lab at Purdue U

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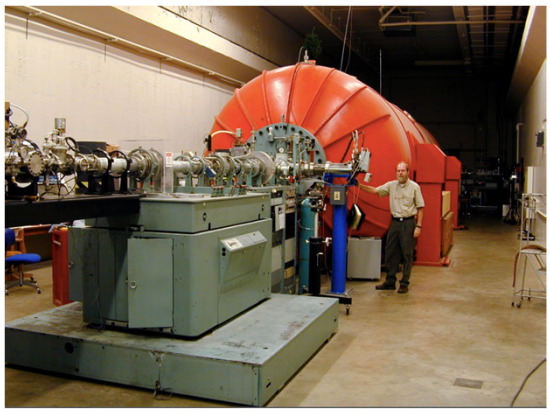
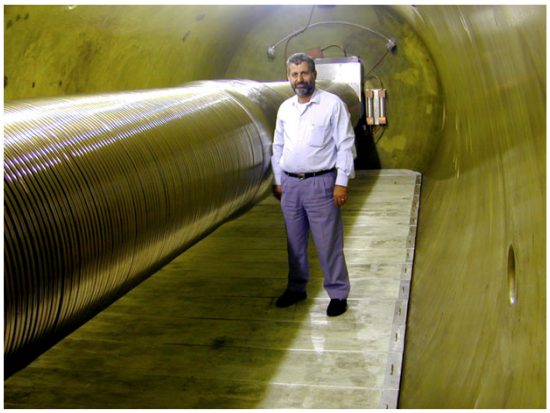
The Van der Graaf accelerator – PRIME lab





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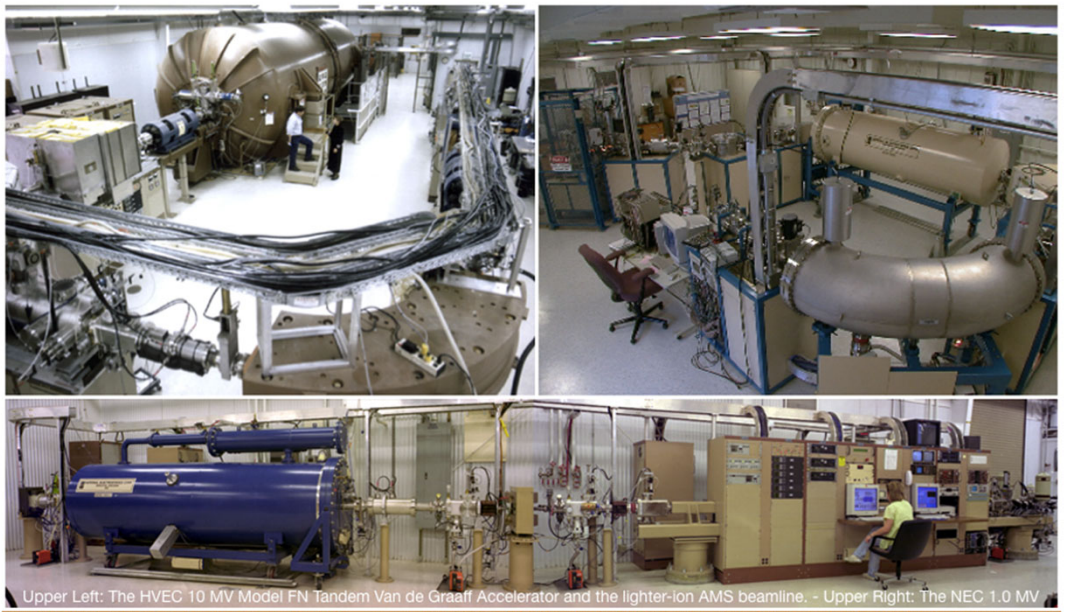

Upper Left: The HVEC 10 MV Model FN Tandem Van de Graaff Accelerator and the lighter-ion AMS beamline. - Upper Right: The NEC 1.0 MV Model 3SDH-1 Tandem Accelerator for BioAMS. - Lower: The NEC 1.7 MV Model 5SDH-2 Tandem Accelerator for Ion Beam Analysis and neutron production.

The AMS facility at the Lawrence Livermore National Laboratory

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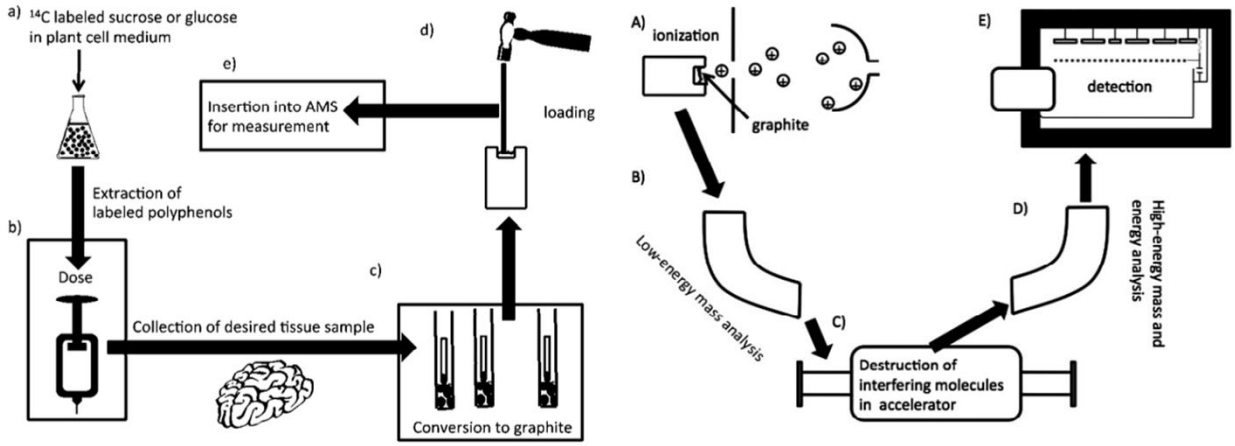
Tracing the appearance of a ^{14}C-labeled precursor



Barnes et al., Food Funct 2011

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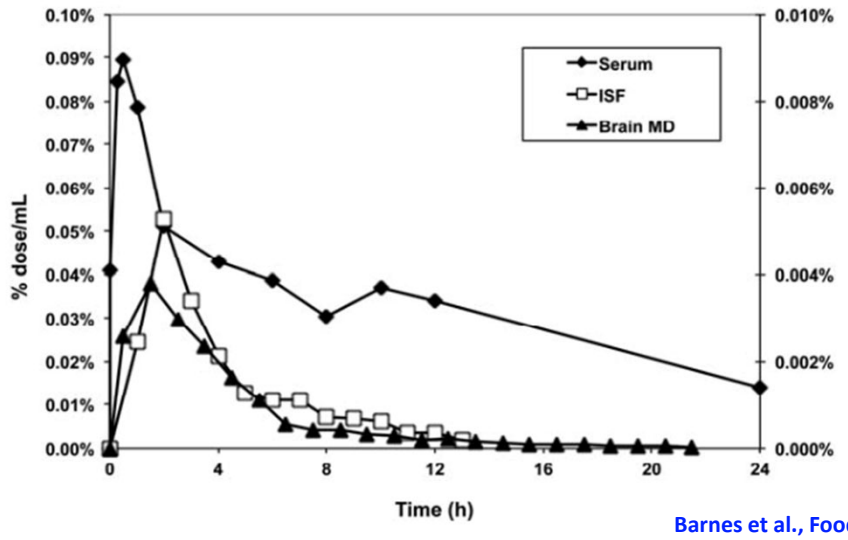
Tracing the movement of ^{14}C-intermediate in tissues



Barnes et al., Food Funct 2011

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Using chemical reagents in metabolomics

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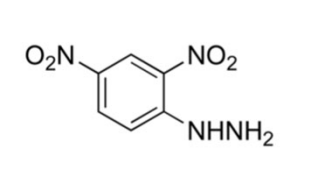
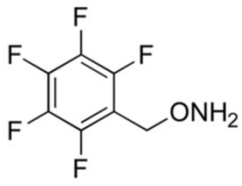
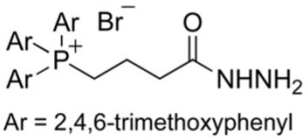
Carbonyl derivatization reagents



DNPH

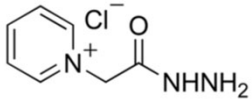


PFBHA

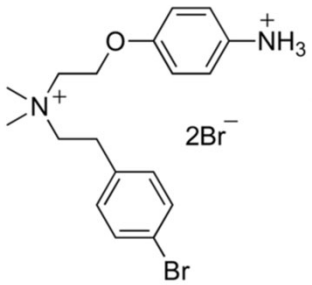


Ar = 2,4,6-trimethoxyphenyl

TMPP-PrG



Girard-P reagent



4-APEBA

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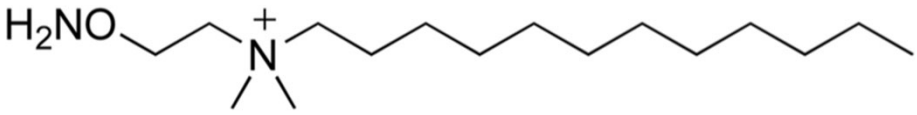
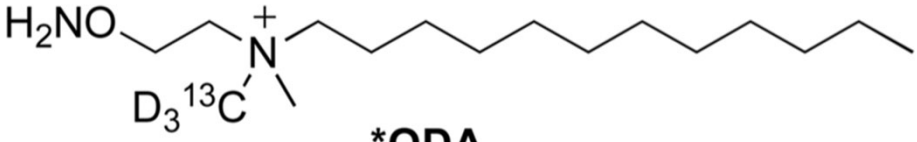
Isotopic carbonyl reagents



QDA



*QDA

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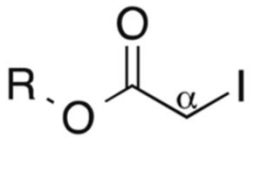
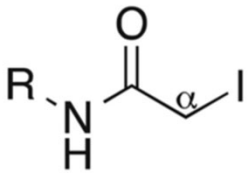
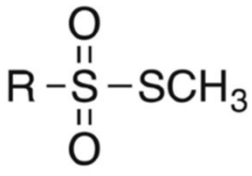
Thiol derivatization reagents



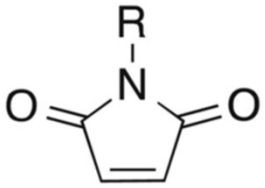
IAA



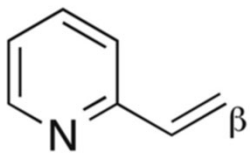
IAM



R = CH_3, **MMTS**



R = CH_3CH_2, **NEM**



VP

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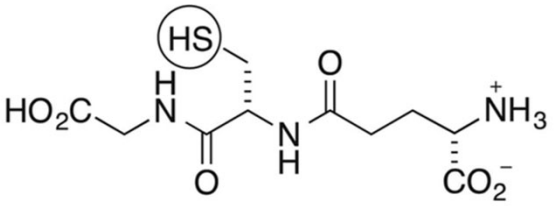
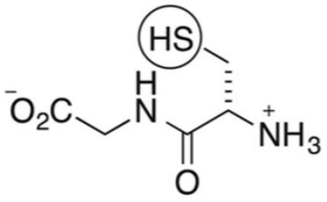
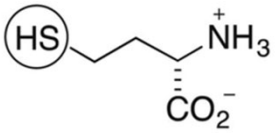
Detectable thio-metabolites



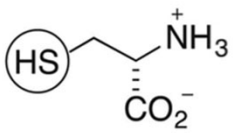
L-glutathione



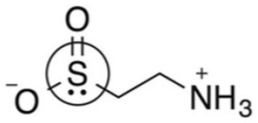
L-cysteinylglycine



L-homocysteine



L-cysteine



hypotaurine

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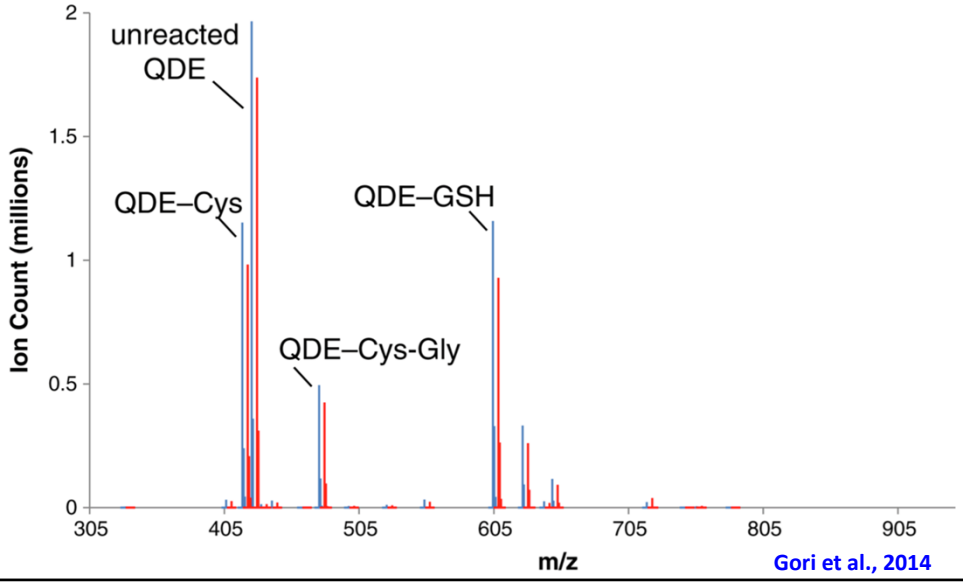
Thiol metabolites in A459 cell extract



Gori et al., 2014

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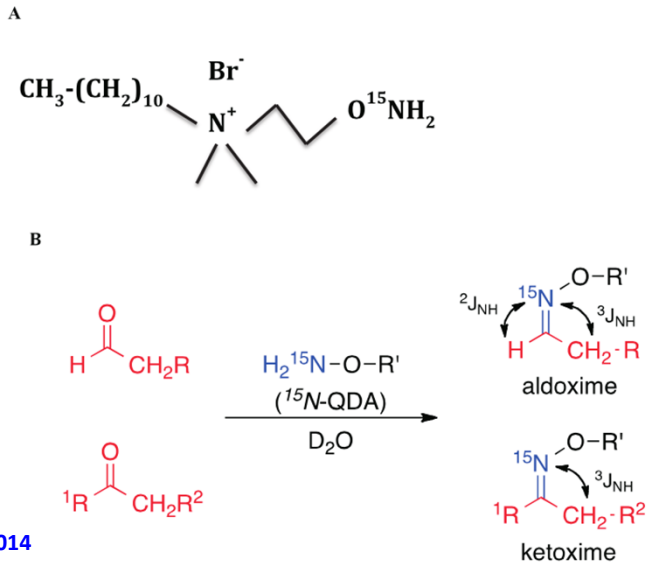
^{15}N-labeled derivatization reagent



Lane et al., 2014

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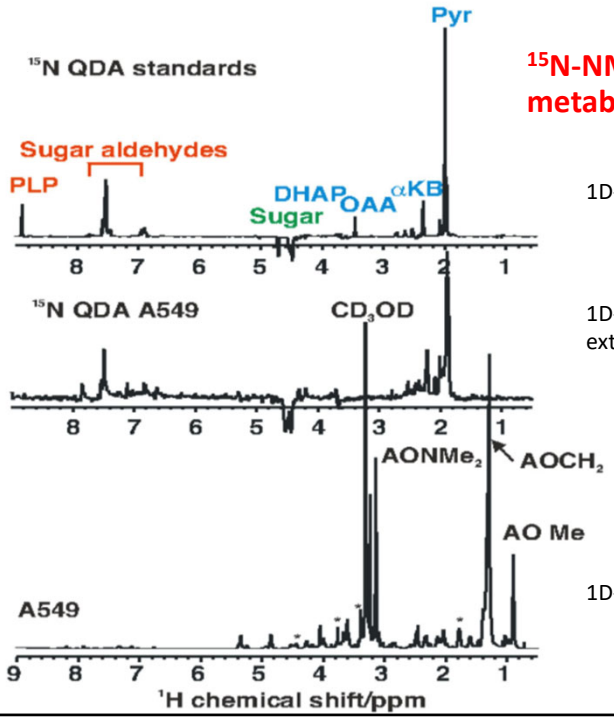

^{15}N-NMR of derivatized metabolites

1D-HSQC spectrum of stds

1D-HSQC spectrum of polar extract of A549 cells

1D-presat spectrum of A549 cells

Lane et al., 2014

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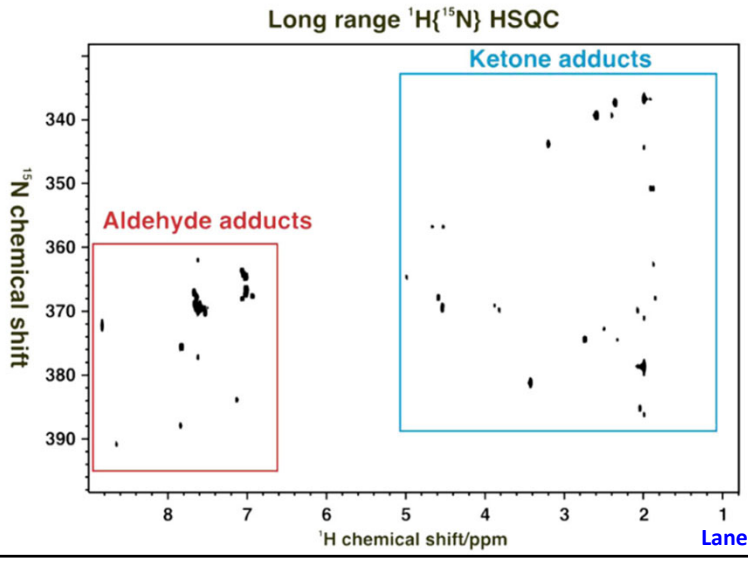
2D-^{1}H, ^{15}N-NMR of standards



Lane et al., 2014

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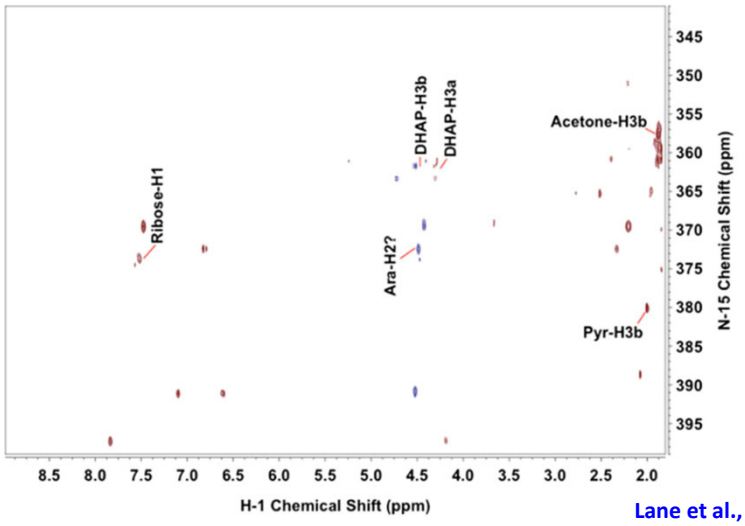
2D-^{1}H, ^{15}N-NMR of A459 cell extract



Lane et al., 2014

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Hyperpolarized NMR

- NMR's advantages
  - Non-destructive
  - Quantitative
  - Usable *in situ*
- NMR's disadvantages
  - The difference in the populations of the low and high energy states is 1 part in a million, i.e., low sensitivity
  - If an excess in the high energy state could be achieved, NMR could become >10,000 times more sensitive, albeit that the half-life in the hyperpolarized state may be only 30 s

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Article

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Kinetic Analysis of Hepatic Metabolism Using Hyperpolarized Dihydroxyacetone

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Making hyperpolarized dihydroxyacetone

- An 8.0 M solution of [2-^{13}C]DHA in a (2:1) water:dimethyl sulfoxide (DMSO) mixture was doped with 15 mM stable trityl free radical (Oxford Instruments Molecular Biotools) and 1.0 mM ProHance.
- The frozen sample was cooled to **1.05 K** in a pumped helium bath inside the magnetic field (3.35 T) of the HyperSense, and the microwave irradiation was turned on.
- Polarization took 1.5–2 h, the irradiation was turned off, and the sample was rapidly dissolved with 4 ml of hot (190 °C) PBS (10 mM, pH 7.4) and transferred to an 89-mm vertical 9.4 T NMR spectrometer for transfer into the perfusate chamber and spectral acquisition.

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Animal experiment and data collection

- Isolated perfused mouse liver placed inside the wide bore of the NMR
- Liver perfused with either 0.2 mM octanoate or 0.2 mM octanoate/2 mM pyruvate
- Hyperpolarized DHA added and first free induction decay collected within 1 msec
- Subsequent FID collection was for 1 sec with a 2 sec delay before the next acquisition – 3 sec cycle time
- Total number of collections = 60, i.e., a total of 3 min
- Half-life of the ^{13}C signal was 32 sec, i.e., by 3 min the signal was 1.5625% of the starting signal

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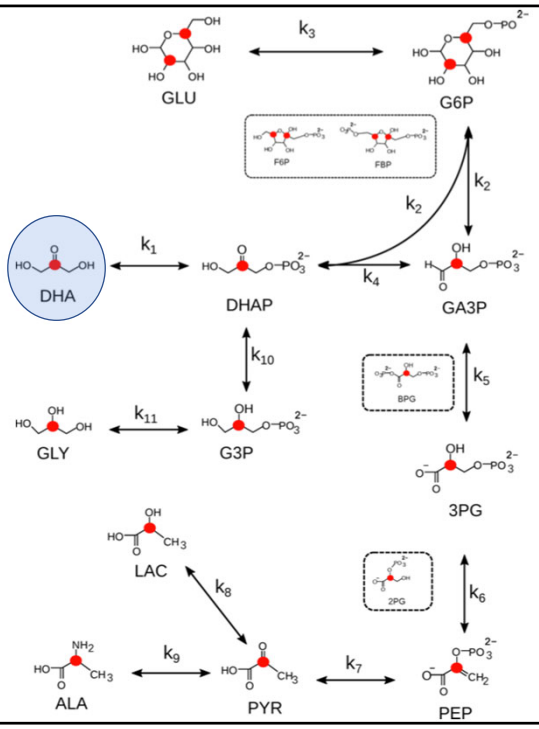

- DHA is first converted to DHA-PO_4
- DHA-PO_4 has a divergence of paths
  - To glycerol-3-PO_4 and then glycerol
  - With GA3P to FDP, G6P and glucose
  - Or by isomerizing to GA3P and then to 3PG, PEP, pyruvate, etc.
- Each step is governed by a linear rate constant, k_1, , k_2, etc.

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Cite this: *Phys. Chem. Chem. Phys.*, 2019, **21**, 22849

Production of highly concentrated and hyperpolarized metabolites within seconds in high and low magnetic fields†

Sergey Korchak,[ab] Meike Emondts,[cd] Salvatore Mamone,[ab] Bernhard Blümich[d] and Stefan Glöggler[*ab]

Hyperpolarized metabolites are very attractive contrast agents for *in vivo* magnetic resonance imaging studies enabling early diagnosis of cancer, for example. Real-time production of concentrated solutions of metabolites is a desired goal that will enable new applications such as the continuous investigation of metabolic changes. To this end, we are introducing two NMR experiments that allow us to deliver high levels of polarization at high concentrations (50 mM) of an acetate precursor (55% ^{13}C polarization) and acetate (17% ^{13}C polarization) utilizing 83% *para*-state enriched hydrogen within seconds at high magnetic field (7 T). Furthermore, we have translated these experiments to a portable low-field spectrometer with a permanent magnet operating at 1 T. The presented developments pave the way for a rapid and affordable production of hyperpolarized metabolites that can be implemented in *e.g.* metabolomics labs and for medical diagnosis.

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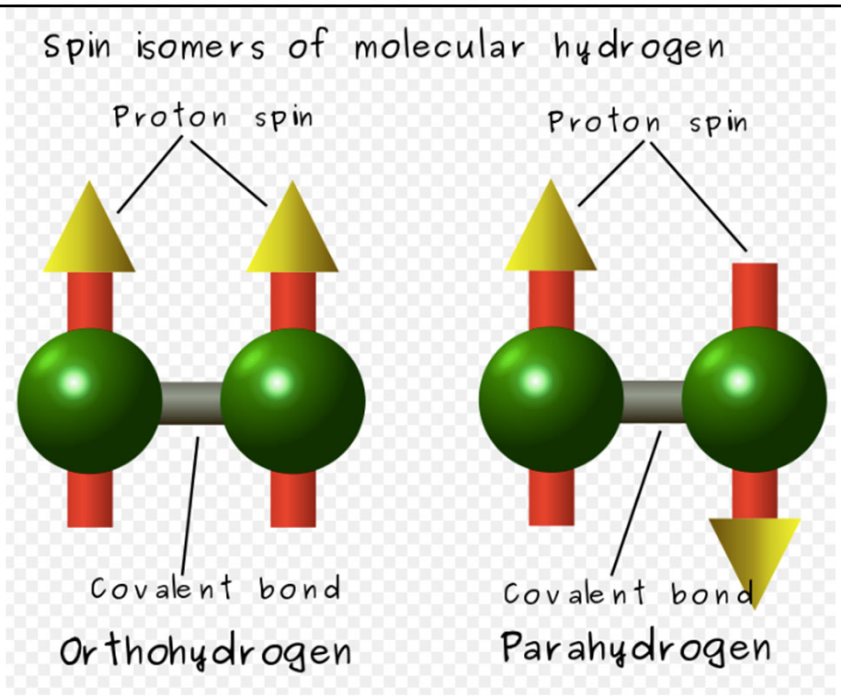

At room temperature, H_2 is 75% ortho and 25% para. Para is the lower energy state

At 1°K, H_2 is predominantly para.

Para H_2 aligns better with a magnetic field and efficiently transfers the polarization to another molecule

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